

Low-temperature catalytic combustion of methane over $\text{MnO}_x\text{--CeO}_2$ mixed oxide catalysts: Effect of preparation method

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The effect of preparation method on $\text{MnO}_x\text{--CeO}_2$ mixed oxide catalysts for methane combustion at low temperature was investigated by means of BET, XRD, XPS, H_2 -TPR techniques and methane oxidation reaction. The catalysts were prepared by the conventional coprecipitation, plasma and modified coprecipitation methods, respectively. It was found that the catalyst prepared by modified coprecipitation was the most active, over which methane conversion reached 90% at a temperature as low as 390 °C. The XRD results showed the preparation methods had no effect on the solid solution structure of $\text{MnO}_x\text{--CeO}_2$ catalysts. More Mn^{4+} and richer lattice oxygen were found on the surface of the modified coprecipitation prepared catalyst with the help of XPS analysis, and its reduction and BET surface area were remarkably promoted. These factors could be responsible for its higher activity for methane combustion at low temperature.

KEY WORDS: $\text{MnO}_x\text{--CeO}_2$ mixed oxide; solid solution; methane combustion; low-temperature activity.

1. Introduction

For the environment and energy consideration, complete oxidation of methane into harmless CO_2 and H_2O has been paid much attention in catalytic combustion field. Among the heterogeneous catalysts, the supported noble metal Pd-based catalysts show excellent activity at low temperature [1–3]. The temperature corresponding to 90% methane conversion (T_{90}) over Pd/ SnO_2 catalyst is 440 °C [2]. Pd/ $\text{Sn}_x\text{Zr}_{1-x}\text{O}_2$ catalyst has been also reported to possess high activity at low temperature for methane complete oxidation, over which T_{90} is 378 °C [3]. However, poor thermal stability and expensive cost of noble metals prevent the widespread application of these catalysts. Recently, the transition metal mixed oxides are attractive due to the relatively low price and as high as or slightly higher catalytic activity toward methane combustion at low temperature than supported noble metals [4,5].

Among the transition metal mixed oxides of interest for oxidation reactions, MnO_x -bases mixed oxide catalysts exhibit great potential. It is generally believed that MnO_x are compounds with a typical berthollide structure and contain labile lattice oxygen. Their catalytic properties are attributed to the ability for manganese to form oxides with variable oxidation states (MnO_2 , Mn_2O_3 , Mn_3O_4 , or MnO) and to their oxygen storage capacity in the crystalline lattice [6,7]. For methane oxidation reaction, the Mn^{4+} sites are stronger catalytic

active sites than the Mn^{3+} sites [6,8]. The active site on supported MnO_x is mainly identified as being Mn^{4+} of MnO_2 in other oxidation processes [7,9]. In addition, CeO_2 has been widely used as a promoter and an oxidation catalyst owing to its unique redox properties and high oxygen storage capacity [10,11]. Compared with pure MnO_x and CeO_2 , $\text{MnO}_x\text{--CeO}_2$ mixed oxides showed higher catalytic activities because manganese and cerium oxides formed solid solution in which oxygen reservoir of CeO_2 and the mobility of oxygen species were greatly enhanced [12,13].

The catalytic performance of $\text{MnO}_x\text{--CeO}_2$ mixed oxides is notably affected by the preparation methods and conditions. The $\text{MnO}_x\text{--CeO}_2$ catalyst modified with Ag exhibited much higher activity at low temperature for oxidation of formaldehyde because the content of lattice oxygen was apparently increased [14]. Previous studies have proved that the modified coprecipitation prepared $\text{MnO}_x\text{--CeO}_2$ catalyst was more active than those prepared by coprecipitation and sol-gel methods [13]. Considering the same catalytic mechanism (Mars-Van-Krevelen redox mechanism) as total oxidation of formaldehyde, modified coprecipitation method was used to prepare $\text{MnO}_x\text{--CeO}_2$ mixed oxide catalysts for methane combustion in this contribution. Meanwhile, as effective molecule activation and surface modification approaches, the plasma technique was also considered to prepare $\text{MnO}_x\text{--CeO}_2$ catalysts in this work. It was reported that the plasma prepared catalysts showed higher dispersion and enrichment on surface of active components, decrease of reduction temperature, etc.

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[15–17]. Since these favorable effects, the low-temperature activities of catalysts prepared by plasma method could be remarkably improved.

With the aforementioned background, the aim of the present work was to investigate the effect of preparation method on the properties of $\text{MnO}_x\text{-CeO}_2$ mixed oxide catalysts for methane combustion at low temperature. $\text{MnO}_x\text{-CeO}_2$ catalysts were prepared by coprecipitation, plasma and modified coprecipitation methods, respectively, and were characterized with BET, XRD, XPS, and TPR techniques.

2. Experimental procedure

2.1. Catalyst preparation

$\text{MnO}_x\text{-CeO}_2$ mixed oxides ($\text{Mn}/(\text{Mn} + \text{Ce}) = 0.5$, molar ratio) were prepared by three different methods. (1) Coprecipitation method: 50% $\text{Mn}(\text{NO}_3)_2$ solution and $(\text{NH}_4)_2\text{Ce}(\text{NO}_3)_6$ with a molar ratio of 5:5 was mixed in dissolved distilled water. Total of 2 M NaOH solution was added to the mixing solution at 50 °C drop by drop until the pH value reached 10.5 with stirring. The mixtures were further aged at 50 °C for 2 h in the mother liquid. After filtration and washing with distilled water, the obtained solid was dried at 110 °C overnight and calcined at 500 °C for 6 h in air. The catalyst was designed as CP. (2) Plasma method: the catalyst was prepared using the same process as the above. Furthermore, the catalyst precursor was treated 90 min by glow discharge plasma before calcination. The treatment approach was similar to that in Ref. [15]. The precursor was put into the discharging tube and decomposed in vacuum. The discharge parameters are as follows: frequency 13.56 MHz, discharge voltage 100 V, anodic current 100 mA. The prepared catalyst was denoted as PP. (3) Modified coprecipitation method: the coprecipitation procedure was the same as CP, but the metal oxide precursors contained KMnO_4 . The molar ratio of $\text{Mn}(\text{NO}_3)_2$, KMnO_4 and $(\text{NH}_4)_2\text{Ce}(\text{NO}_3)_6$ was 1:4:5. The catalyst was nominated as MP. The precipitates were filtered and washed with distilled water many times during the preparation as Ref [13]. There should be no or considerable little potassium in the MP catalyst and the effect of traces of potassium on catalytic activity was not considered in this work. For a comparison purpose, pure MnO_x was prepared by coprecipitation of $\text{Mn}(\text{NO}_3)_2$ and KMnO_4 with the ratio of 1:4.

2.2. Catalyst characterization

The BET surface area of catalysts were determined by N_2 adsorption experiments at -196 °C on a micromeritics NOVA 1000 e. Before each measurement, the sample was degassed in vacuum at 300 °C for 3 h.

X-ray powder diffraction (XRD) patterns were recorded on the DX-2000 diffractometer using $\text{Cu K}\alpha$

($\lambda = 1.54056$ Å) radiation between 10° and 80°. The voltage and anode current were 40 kV and 30 mA, respectively. The average crystallite sizes of the cubic phase were evaluated from the half-width of the ceria (1 1 1) peak according to the Scherrer's equation [18]. The lattice parameters were calculated according to the full profile fitting procedure.

X-ray photoelectron spectra (XPS) experiments were performed on the XSAM800 spectrometer with a Al anode for $\text{K}\alpha$ (1486.6 eV) radiation. Charging effects were corrected by adjusting the binding energy of C1s peak from carbon contamination to 284.6 eV.

Temperature-programmed reduction (TPR) measurements were carried out at atmospheric pressure in a fixed-bed reactor. Fifty milligrams sample was loaded and pretreated with N_2 at 300 °C for 1 h to remove the adsorbed carbonates and hydrates. After cooling down to 50 °C, the reduction agent of 4.22% H_2/N_2 with a flow rate of 30 mL/min was introduced. The temperature of the reactor was increased linearly from 50 to 650 °C at a rate of 10 °C/min by a temperature controller. The effluent stream was analyzed for hydrogen by a thermal conductivity detector.

2.3. Catalytic activity test

Methane combustion was performed in a fixed-bed reactor with a continuous flow at atmospheric pressure in the temperature range of 300–480 °C. In order to decrease the effect of exothermic reaction on bed-temperature, catalyst sample of 150 mg was diluted in 300 mg silica sand. Subsequently, the diluted catalyst was packed in a quartz tube that connected with a thermocouple centered in the catalyst bed. The reaction gases consisted of 20% CH_4 and 40% O_2 in Ar, and space velocity (SV) was 40,000 $\text{ml g}^{-1} \text{h}^{-1}$. The analysis of the effluent from the reactor was performed by the 112A model gas chromatograph with a thermal conductivity detector (TCD).

3. Results and discussion

3.1. Catalyst activity in methane combustion

Figure 1 illustrates the results of activities of various catalysts for methane combustion. In this work, CO_2 and H_2O are the only reaction products of methane oxidation. As expected, the preparation methods significantly affect the catalytic activities of $\text{MnO}_x\text{-CeO}_2$ mixed oxides. For the CP, PP and MP catalysts, the temperatures responding to 50% CH_4 conversion (T_{50}) were 408, 376 and 348 °C, respectively. The T_{50} of the PP and MP catalysts decreased 32 and 60 °C in contrast to that of the CP catalyst, indicating that the MP and PP catalysts possessed much higher activity in the experimental temperature range, especially the MP catalyst. Moreover, the T_{90} over the MP catalyst was as low as

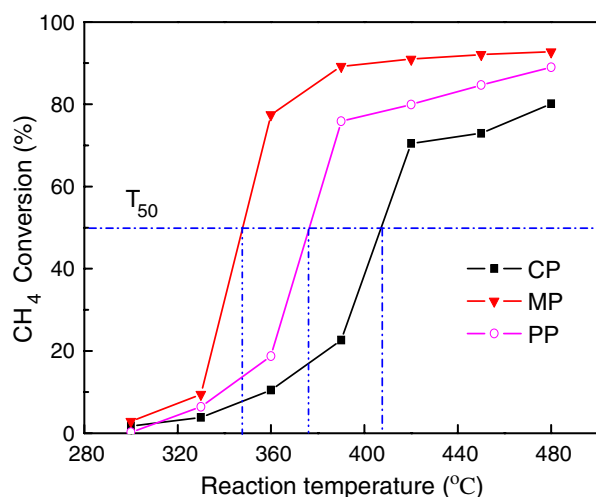


Figure 1. Catalytic activity in methane total oxidation (Reaction conditions: 150 mg catalyst, $V(\text{CH}_4):V(\text{O}_2):V(\text{Ar}) = 20:40:40$, $\text{SV} = 40,000 \text{ ml g}^{-1} \text{ h}^{-1}$).

390 °C. When the specific conversion (defined as methane conversion per surface area unit) was considered, the activity difference was still obvious below 390 °C. Moreover, the rates of methane combustion ($\text{mol/m}^2 \text{ s}$) on the three catalysts were compared, and similar results were also obtained. Thereby, it could be claimed that the MP catalyst was the most active at low temperature.

3.2. BET texture analysis

The surface area, pore volume and average pore diameter of $\text{MnO}_x\text{-CeO}_2$ mixed oxide catalysts are summarized in table 1. From table 1, it can be seen that the surface area and pore volume of the MP catalyst ($106.40 \text{ m}^2/\text{g}$ and $0.214 \text{ cm}^3/\text{g}$, respectively), are much larger than those of the CP catalyst. In addition, the surface area of the PP catalyst ($80.11 \text{ m}^2/\text{g}$) slightly increases in comparison with that of the CP catalyst ($74.59 \text{ m}^2/\text{g}$), while the average pore diameter decreases by 45%. All the average pore diameters of the catalysts are below 10 nm. The results of BET measurements suggest that the surface areas of $\text{MnO}_x\text{-CeO}_2$ mixed oxides are related to their preparation methods, and the MP catalyst shows a maximum surface area.

3.3. XRD characterization

Figure 2 displays the XRD patterns of $\text{MnO}_x\text{-CeO}_2$ catalysts. There was no diffraction of manganese oxides

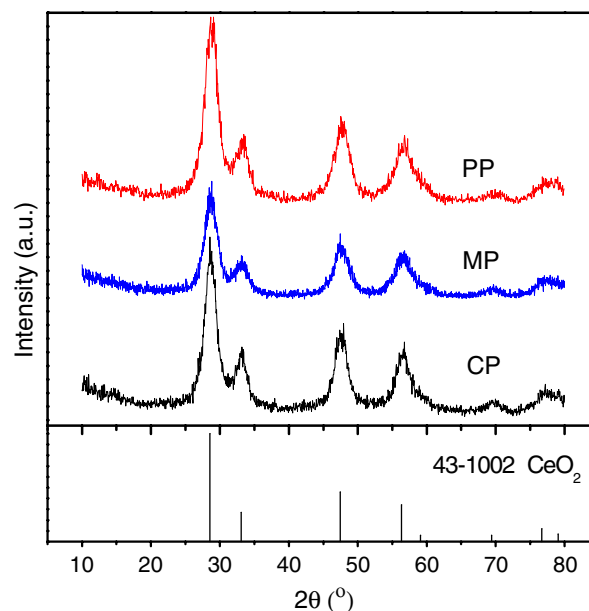


Figure 2. XRD patterns of $\text{MnO}_x\text{-CeO}_2$ catalysts.

in the figure. Only the broad peaks of cubic fluorite structure CeO_2 (JCPDS #43–1002) were observed. The XRD qualitative analysis results show that $\text{MnO}_x\text{-CeO}_2$ mixed oxides crystallize in CeO_2 type structure as reported in Ref [19]. Mn atoms probably replace some of Ce sites and enter the CeO_2 lattice.

The refined lattice parameters and crystallite sizes of the $\text{MnO}_x\text{-CeO}_2$ mixed oxide catalysts are listed in table 2. As the ion radius of Mn ($r(\text{Mn}^{3+}) = 0.66 \text{ Å}$ and $r(\text{Mn}^{4+}) = 0.60 \text{ Å}$) is smaller than the ion radius of Ce ($r(\text{Ce}^{4+}) = 0.92 \text{ Å}$), and Mn can enter CeO_2 lattice, the lattice parameters of these mixed oxides are smaller than that of Pure CeO_2 (5.411 Å). The XRD results identify that the preparation methods have no effect on the phase composition of $\text{MnO}_x\text{-CeO}_2$ mixed oxides. In addition, it can be also seen that the average crystallite sizes of the CP, PP and MP catalysts are 49, 29 and 42 Å, respectively. Obviously, the crystallite size of the PP catalyst decreased notably, indicating that the glow discharge plasma treatment significantly enhances the dispersion of active components.

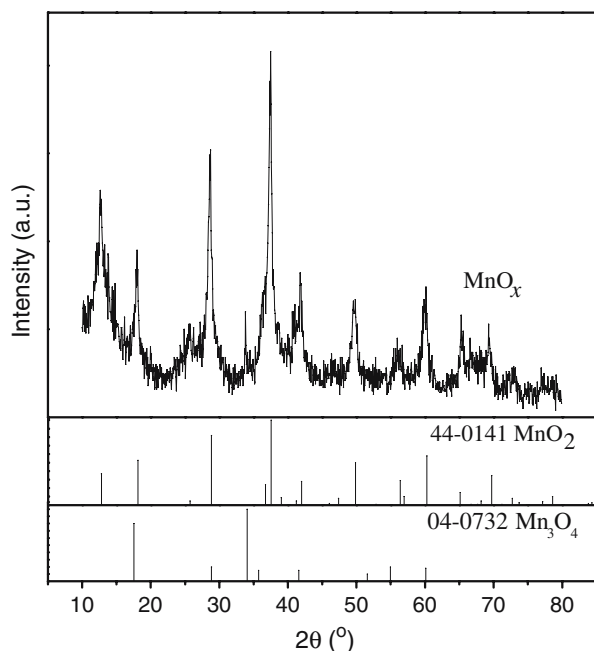
According to XRD patterns showed in figure 3, MnO_2 (JCPDS #44–0141) was the main phase of the pure MnO_x with a much lower content Mn_3O_4 (JCPDS #04–0732). Consequently, it might be accepted that manganese species was almost Mn^{4+} in the MP catalyst.

Table 1
Surface area, pore volume, average pore diameter of various catalysts

Catalyst	Surface area (m^2/g)	Pore volume (cm^3/g)	Average pore diameter (nm)
CP	74.5(9)	0.17(9)	9.61(2)
PP	80.1(1)	0.13(7)	6.84(1)
MP	106.4(0)	0.21(4)	8.03(7)

Table 2
Lattice parameters and crystallite sizes of $\text{MnO}_x\text{-CeO}_2$ catalysts

Catalyst	Lattice parameter (Å)	Crystallite size (Å)
CP	5.403 (0)	49 (2)
PP	5.382 (2)	29 (3)
MP	5.389 (4)	42 (2)

Figure 3. XRD patterns of pure MnO_x sample.

3.4. XPS analysis

The XPS spectra of Mn2p, O1s and Ce3d were measured for these three $\text{MnO}_x\text{-CeO}_2$ catalysts. From the XPS spectra of Mn2p, for the CP catalyst, the binding energies around 643.49s and 641.86 eV could be ascribed to the presence of Mn^{4+} and Mn^{3+} species, respectively. Similarly, the binding energies around 643.00 and 641.23 eV was also attribute to Mn^{4+} and Mn^{3+} species for the PP catalyst. In the MP catalyst, it is interesting to note that only one peak at about 642.55 eV corresponding to Mn^{4+} species was observed within the resolution of XPS. This result is in good consistent with the XRD measurements. According to the corresponding XPS spectra of O1s, two peaks (O_I and O_{II}) were displayed, which represent two different kinds of surface oxygen species. O_I with BE from 529.15 to 529.31 eV is characteristic of the lattice oxygen (O^{2-}) [19,20], while O_{II} with BE of 530.22–531.65 eV belongs probably to the defect oxides or the surface oxygen ions with low coordination [13,14]. From the XPS spectra of Ce3d, the binding energies of Ce^{4+} were located at the characteristic values of 882.40–882.60 eV as reported in Ref. [21], suggesting that the oxidation state of cerium

was CeO_2 in all the catalysts. This is also in agreement with the XRD results.

The surface compositions calculated from the XPS spectra of the $\text{MnO}_x\text{-CeO}_2$ catalysts are summarized in table 3. The different preparation methods resulted in remarkable varieties of relative concentration of Mn^{4+} and O_I . In the CP catalyst, the relative concentrations of Mn^{4+} was calculated to be about 30%, but the data in the PP and MP samples were almost 80% and 100%, respectively. The corresponding concentrations concentrations of O_I were 13.63%, 36.58% and 53.77%, respectively. This reveals that plasma and modified coprecipitation methods can apparently increase concentrations of Mn^{4+} and lattice oxygen species on the surface in contrast to conventional coprecipitation method, especially modified coprecipitation method. Therefore, the MP sample possessed the most Mn^{4+} and richest lattice oxygen species. The surface condition of the MP catalyst obviously favors the oxidation of methane, and provides a good interpretation for the higher activity of the MP catalyst at low temperature.

3.5. $\text{H}_2\text{-TPR}$ measurements

The $\text{H}_2\text{-TPR}$ profiles of the pure MnO_x and the $\text{MnO}_x\text{-CeO}_2$ mixed oxides are showed in figure 4. From the figure 4, $\text{H}_2\text{-TPR}$ profiles of the pure MnO_x showed an intensive reduction peak with a maximum at about 370 °C, which could be attributed to the reduction of $\text{MnO}_2\text{-MnO}$. Since the larger negative value of reduction potential, the reduction of MnO to Mn^0 has not been observed even up to 950 °C [22]. Therefore, MnO was regarded as the final reduction state.

Compared with the reduction behaviors of pure MnO_x , the reduction temperatures of the $\text{MnO}_x\text{-CeO}_2$ mixed oxides systematically shifted to lower regions. This indicated that the reductions of $\text{MnO}_x\text{-CeO}_2$ mixed oxides were promoted due to the formation of the solid solution. However, the reduction behaviors of the catalysts are significantly influenced by the preparation methods as showed in figure 4. The $\text{H}_2\text{-TPR}$ profile of the CP catalyst exhibited an intensive reduction peak with a maximum at 346 °C, followed by a weak reduction peak at about 430 °C. The two reduction peaks could be attributed to the reductions of Mn_2O_3 to MnO and surface CeO_2 , respectively. Comparatively, the reduction of the PP catalyst began in much lower region

Table 3
XPS results of the $\text{MnO}_x\text{-CeO}_2$ catalysts

Catalyst	BE (eV)		$\text{Mn}^{4+}/(\text{Mn}^{4+} + \text{Mn}^{3+})(\%)$	BE (eV)		$\text{O}_I/(\text{O}_I + \text{O}_{II})(\%)$	BE (eV)
	Mn^{4+}	Mn^{3+}		O_I	O_{II}		$\text{Ce3d}_{5/2}$
CP	643.4(9)	641.8(6)	32.7(9)	529.1(7)	530.2(2)	13.6(3)	882.5(3)
PP	643.0(0)	641.2(3)	78.9(7)	529.3(1)	531.6(5)	36.5(8)	882.4(9)
MP	642.5(5)	—	100	529.1(5)	531.1(3)	53.7(7)	882.5(3)

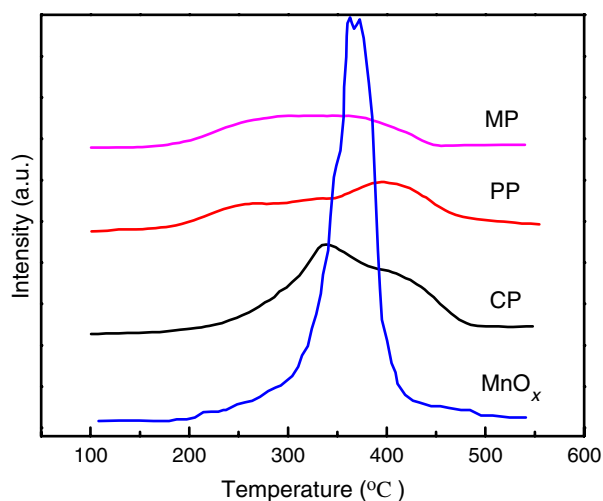
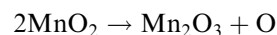


Figure 4. H₂-TPR profiles of MnO_x and MnO_x-CeO₂ catalysts.

(at a maximum of about 250 °C). The low temperature reduction could be ascribed to the reduction MnO₂/Mn₂O₃ to Mn₃O₄, and the high temperature reduction assigned to the combined reduction of Mn₃O₄ to MnO and surface oxygen removal of CeO₂. For the MP catalyst, the H₂-TPR profile exhibited overlapped strong reduction peaks from 170 to 450 °C. According to fitting multi-peaks (Gauss model) with origin 7.0 software, it was found that the overlapped reduction peaks consisted of three peaks with the maximums at about 276, 356, and 400 °C, respectively. In general, the reduction profile of MnO₂ shows a typical two-step reduction process, and the reduction sequence is probably MnO₂ → Mn₃O₄ → MnO [13,23]. However, Craciun et al. [6] presented a three-step reduction profile for unsupported MnO_x. The first two peaks at about 327 and 417 °C correspond to two-step reduction of MnO₂ (MnO₂ → Mn₂O₃ → Mn₃O₄). The third peak at 510 °C represents the total reduction of Mn₃O₄ to MnO. Three reduction temperatures of the MP catalyst obviously shifted to lower regions compared with those of unsupported MnO_x [6]. The solid solution between MnO_x and CeO₂ was formed in which mobility of oxygen species was greatly promoted. As a result, the reduction temperature of MnO₂ in the MP catalyst was decreased significantly.

Based on the results of characterization and reaction, there were more MnO₂ and lattice oxygen and larger surface area for the modified coprecipitation prepared catalyst, resulting in higher activity toward methane combustion at low temperature. As showed in XRD patterns, MnO₂ and CeO₂ in the MP catalyst formed the solid solution, leading to strongly synergistic interaction. According to three reactions represented by Ding et al. [24], the synergistic mechanism could be explained by considering the effective activation of molecule oxygen over the MnO₂-CeO₂ solid solution.



4. Conclusion

MnO_x-CeO₂ mixed oxide catalysts are highly active for methane combustion at low temperature. The catalytic activity primarily depended on the preparation methods. The catalyst prepared by modified coprecipitation exhibited much higher activity toward methane oxidation than those prepared by coprecipitation and plasma methods, and the methane conversion reached 90% at a temperature as low as 390 °C over the MP catalyst. XRD characterization identified that the preparation methods had no effect on the phase formation of MnO_x-CeO₂ catalysts. However, BET, XPS and H₂-TPR measurements suggested that larger surface area, more Mn⁴⁺, richer lattice oxygen and easier reduction resulted in higher activity of the MP catalyst.

Acknowledgments

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References

- [1] P. Gélin and M. Primet, *Appl. Catal. B* 39 (2002) 1.
- [2] K. Sekizawa, H. Widjaja and S. Maeda, *Catal. Today* 59 (2000) 69.
- [3] W. Lin, L. Lin and Y.X. Zhu, *Appl. Catal. B* 57 (2005) 175.
- [4] C.-J. Zhou, W. Lin and Y.-X. Zhu, *Chinese J. Catal.* 24 (2003) 229.
- [5] L.F. Liotta, G. Di Carlo, and G. Pantaleo, et al. (2005) *Appl. Catal. B* (doi:10.1016/j.apcatb.12.023).
- [6] R. Craciun, B. Nentwick and K. Hadjiivanov, *Appl. Catal. A* 243 (2003) 67.
- [7] R. Craciun, *Catal. Lett.* 55 (1998) 25.
- [8] A. Machocki, T. Ioannides and B. Stasinska, *J. Catal.* 227 (2004) 282.
- [9] K. Hadjiivanov and J.-C. Lavalley, *Catal. Commun.* 2 (2001) 129.
- [10] K. Otsuka, Y. Wang and M. Nakamura, *Appl. Catal. A* 183 (1999) 317.
- [11] A. Trovarelli, C. de Leitenburg and M. Boaro, *Catal. Today* 50 (1999) 353.
- [12] A.M.T. Silva, R.R.N. Marques and R.M. Quinta-Ferreira, *Appl. Catal. B* 47 (2004) 269.
- [13] X. Tang, Y. Li and X. Huang, *Appl. Catal. B* 62 (2006) 265.
- [14] X. Tang, J. Chen and Y. Li, *Chem. Eng. J.* 118 (2006) 119.
- [15] M.H. Chen, W. Chu and X.Y. Dai, *Catal. Today* 89 (2004) 201.
- [16] Y.-P. Zhang, P.-S. Ma and X. Zhu, *Catal. Commun.* 5 (2004) 35.

- [17] J.-g. Wang, C.-j. Liu and Y.-p. Zhang, , Catal. Today 89 (2004) 183.
- [18] A.L. Patterson, Phys. Rev. 56 (1939) 978.
- [19] M. Machida, M. Uto and D. Kurogi, , Chem. Mater. 12 (2000) 3158.
- [20] H. Chen, A. Sayari and A. Adnot, , Appl. Catal. B 32 (2001) 195.
- [21] F. Larachi, J. Pierre and A. Adnot, , Appl. Surf. Sci. 195 (2002) 236.
- [22] J. Carno, M. Ferrandon and E. Bjornbom, , Appl. Catal. A 155 (1997) 265.
- [23] M.C. Alvarez-Galvan, V.A. de la Pena O'Shea, J.L.G. Fierro and P.L. Arias, Catal. Commun. 4 (2003) 223.
- [24] Z.Y. Ding, L. Li and D. Wade, , Ind. Eng. Chem. Res. 37 (1998) 1707.